

Blitz, Units 18-23, Form I-L

Name _____ Period _____

This is a Take Home Exam. You may use your Notes, PowerPoint, or Text on this exam but NO help from human beings!

You MUST HAND WRITE THIS EXAM!! NO TYPED PAPERS WILL BE ACCEPTED!

EXPLAIN IN COMPLETE SENTENCES AND GIVE EXAMPLES

1. Correct to STP (watch your temperature units!-- 25.7 liters of CH₄ gas at 39.2°K and 156 KPa.
2. Explain these: Solute, solvent, solution, saturated, unstaturated, supersaturated, colloid, tyndall effect.
3. Explain TWO principles that enable us to boil water in a paper cup.
4. Balance the equation for the burning of C₄H₁₀ and write the big K for it.
5. Discuss THREE factors determining the rate of a chemical reaction.
6. Explain the Entropy of dissolving gases into liquids. Give an example from real life.
7. Using $PV = nRT$, solve for the volume in L when the pressure is 152.5 KPa, the temperature is 16°C, and the number of moles is 2.13. (Watch temp. units!).
 $R = 8.31 \text{ L}\cdot\text{KPa}/\text{mol}\cdot^\circ\text{K}$
8. Write the net ionic equation for the Precipitation of Sb(OH)₃.
9. Using graphs and WORDS, Explain the activation energies of endothermic, exothermic, and catalytic reactions.
10. Explain {Spontaneous Combustion}Aaaaaaaaagh..

When finished, please STAPLE this exam onto your papers and turn in on due date.